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## Density Practice Questions - Corbettmaths

To calculate the answer here, we need to recall the formula:  $\text{density} = \frac{\text{mass}}{\text{volume}}$  In this question, the mass is 2460 kg and the volume is  $1.2 \text{ m}^3$ , so we simply need to substitute these values into the formula as follows:  $\text{Density} = 2460 \text{ kg} \div 1.2 \text{ m}^3 = 2050 \text{ kg/m}^3$

## Density Mass Volume Worksheets | Questions and Revision | MME

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## **Answer Key For Density Worksheets - Teacher Worksheets**

Answers 1)  $d = 40.5 \text{ g} / 15.0 \text{ mL}$   $d = 2.7 \text{ g/mL}$  2)  $d = 305.4 \text{ g} / 22.5 \text{ mL}$   $d = 13.57333 \text{ g/mL}$  3)  $d = \text{g} / \text{mL}$   $\text{g} = (d) (\text{mL})$   $d = 2.70 \text{ g/mL}$   $d = 13.6 \text{ g/mL}$   $\text{g} = (0.789 \text{ g/mL}) (200.0 \text{ mL}) = 158 \text{ g}$  4)  $(8.4 \text{ cm}) (5.5 \text{ cm}) (4.6 \text{ cm}) = 212.52 \text{ cm}^3$   $d = 1896 \text{ g} / 212.52 \text{ cm}^3 = 8.9$

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g/cm<sup>3</sup> Significant figures in the answer are dictated by the length measurements of two sig figs. 5) mass of CCl<sub>4</sub> = 703.55 g - 345.8 g = 357.75 g  $d = 357.75 \text{ g} / 225 \text{ mL} = 1.59 \text{ g/mL}$  6)  $d = 65.14 \text{ g} / 35.4 \text{ mL} = 1.84 \text{ g/mL}$  7)  $d = \text{g} / \text{mL}$   $g = (d) (\text{mL} \dots)$

## **density practice problems key - studylib.net**

Density Practice: Worksheet #1. Calculate density, and identify substances using a density chart. Density is a measure of the amount of mass in a certain volume. This physical property is often used to identify and classify substances. It is usually expressed in grams per cubic centimeters, or g/cm<sup>3</sup>.

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40.5 g. What is its density? 2) Mercury metal is poured into a graduated cylinder that holds exactly 22.5 mL. The mercury used to fill the cylinder weighs 306.0 g. From this information, calculate the density of mercury.

### **Density Practice Problem Worksheet - CHSD**

Density Practice Problem Worksheet 1) A block of aluminum occupies a volume of 15.0 mL and weighs 40.5 g. What is its density? 2) Mercury metal is poured into a graduated cylinder that holds exactly 22.5

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## **DSoftSchools**

One way to do this is to mentally compare your answer against the density of water (1 gram per cubic centimeter). Light substances would float on water, so their density should be less than that of water. Heavy materials should have density values greater than that of water.

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$m(\text{silver}) = D \times v$   $m(\text{gold}) = D \times v$ .  $m(\text{silver}) = 10.5$   
 $\text{g/cm}^3 \times 5 \text{ cm}^3$   $m(\text{gold}) = 19.3 \text{ g/cm}^3 \times 5 \text{ cm}^3$ .  $m$   
 $(\text{silver}) = 52.5 \text{ gm}$   $(\text{gold}) = 96.5 \text{ g}$ . Answer gold. 6.

Five mL of ethanol has a mass of 3.9 g, and 5.0 mL of benzene has a mass of 44 g. Which liquid is denser?  $v$   
 $(\text{ethanol}) = 5 \text{ ml}$   $v(\text{benzene}) = 5 \text{ ml}$ .

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## **Density Practice Problems - New Canaan**

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